UNIT 4 TEST 2 MS

1. (a) (i) Tollens' reagent / ammoniacal silver nitrate **(1)**

silver mirror / precipitate **(1)**

butanoate / butanoic acid / unambiguous formula or structure **(1)** 3

(ii) **Any of:**

Br_2 – decolourises – (electrophilic) addition
Na – fizzes – redox
$SOCl_2$ /PCl_5 /acid chloride – white fumes – substitution/chlorination
carboxylic acid + conc H_2SO_4 /acid chloride – sweet smell – esterification/ condensation

test **(1)** - observation **(1)** - type of reaction **(1)** 3

NOT
2-4DNPH to give no precipitate

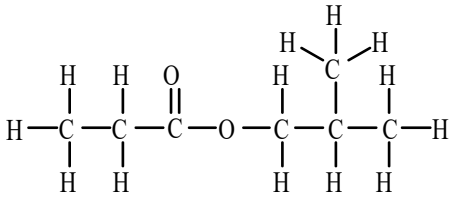
(b) recrystallise /purify (the precipitate) **(1)**

measure melting point **(1)**

compare with known values **(1)** 3

[9]

2. (a)



propanoate and ester group **(1)** 2-methyl propyl **(1)** 2

(b) propanoic acid **(1)**
(2-)methylpropan-1-ol **(1)**

heat **(1)**
conc. H_2SO_4 **(1)**
(allow ecf from part (a) for the equation)

$CH_3CH_2COOH + CH_3)_2CHCH_2OH \rightarrow CH_3CH_2COOCH_2CH(CH_3)_2 + H_2O$

reactants **(1)** products **(1)** 6

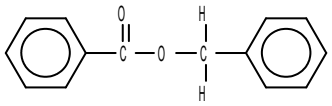
(c) mass spectrum / spectrometry **(1)**

molecular ion peak /
m/e or mass of the peak furthest right **(1) AW** 2

[10]

3. (i) H^+ /acid / named strong acid eg H_2SO_4 / HC*l* 1

(ii)



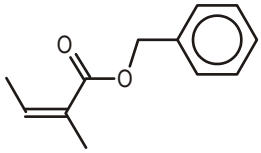
displayed ester group **(1)**
rest of the ester **(1)** 2

[3]

4. (a) (i) alkene **(1)**
ester **(1)** 2

allow "C=C double bond"

i.



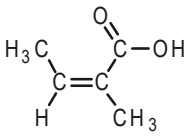
(1) 1

ii. $C_{12}H_{14}O_2$ **(1)** 1

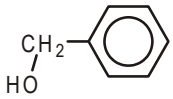
(b) same structural formula/order of bonds,
different spacial arrangement **AW (1)**

description or diagram showing **B** and how it is different from **A (1)** 2

(c)



(1)



(1) 2

(d) (i) peak at 1680-1750 (cm^{-1}) due to C=O **(1)**

peak at 1000-1300 (cm^{-1}) due to C-O / **(1)** 2

(ii) 2500-3300 / 3230-3550 (cm^{-1}) **(1)**

O-H /carboxylic acid/alcohol is **not** present in **A (1)**

allow 1 mark for ~500-1500 (cm^{-1}) which is a unique fingerprint region etc 2

[12]

5. (i) hydrolysis **(1)**

(sorbitan monolaurate is an) ester **(1)**

broken down to form an alcohol and carboxylic acid/salt **(1) AW**
/ equation to show the reaction 3

(ii) sorbitan monolaurate is made from a renewable resource
/ not based on crude oil **(1) AW** 1

[4]

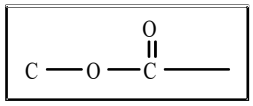
6. Three of following points: **(1)(1)(1)**

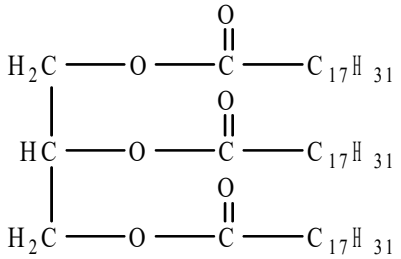
- 1. There is van der Waals (IDID) between triglycerides.
- 2. There is van der Waals between triglycerides and (non-polar) solvent.
- 3.Triglycerides cannot hydrogen bond (to water)(enough).
- Because there are not enough suitable sites/oxygen atoms
  Or long hydrocarbon chains do not hydrogen
  bond/would interfere with hydrogen bonding in water
  AW 3

[3]

7. (i)

(1) for a correct ester
(1) for rest





Accept correct skeletal form (even if only for acyl groups)
but must have 17C and two double bonds/one triple bond 2

(ii) 6. Ecf from (i). **(1)** 1

[3]

8. (i) CH_3COOH **(1)** 1

(iv) CH_3COOH **(1)**
CH_3CH_2OH **(1)** 2

[3]